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The video below refers to Dynamic Equilibrium (under the section Equilibrium Systems) 7.1 p. 420 in the Nelson Textbook The video below is related to 7.2 Equilibrium Law and the Equilibrium Constant (p.429 in the Textbook)

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Solution: OH OH OH 0.50 mmol mL -- -- = x = nc V x 25.0 mL OH n -- = 12.5 mmol HCl(aq) HCl(aq) [HCl(aq)] 12.5 mmol = = n V 0.10 mmol HCl(aq) /mL V =120 mL 2. (a) Given: V HCl(aq) = 25.00 mL; [HCl(aq)] = 0.350 mol/L; V NaOH(aq) = 5.0 mL; [NaOH(aq)] = 0.500 mol/L; Required: [H+(aq)], pH Analysis: Before titration, amount of H+ is: H(aq) H(aq) H(aq) 0.350 mmol mL ++ += x = nc V

Section 8.7: Acid – Base Titration Tutorial 1 Practice, page 547

Complex chemistry concepts are presented in a clear, understandable fashion and key concepts, such as thermodynamics, are treated in greater depth than specified in the curriculum. Nelson Chemistry 12 provides a rigorous, comprehensive, and accurate treatment of all concepts and processes presented in Ontario's chemistry, Grade 12, university Preparation course (SCH4U).

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Solution: $Q = \frac{[H_2(g)][I_2(g)]}{[HI(g)]^2} = \frac{(0.040)(0.010)}{(0.14)^2}$ $Q = 0.020$ Statement: The value of Q is 0.020. Q is equal to K , so the reaction is at equilibrium. (b) Given: $[HI(g)] = 0.20 \text{ mol/L}$; $[H_2(g)] = 0.15 \text{ mol/L}$; $[I_2(g)] = 0.090 \text{ mol/L}$; $K = 0.020$ Required: Q ; direction of reaction to reach equilibrium
Solution: $Q = \frac{[H_2(g)][I_2(g)]}{[HI(g)]^2} = \frac{(0.15)(0.090)}{(0.20)^2}$ $Q = 0.34$

Section 7.5: Quantitative Changes in Equilibrium Systems ...

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