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He wrote the first practical extrusion book back in the 1960s as well as the Plastics Extrusion Operating Manual, updated

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almost every year ... key principles of extrusion, plastics chemistry for ...

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“ The system has already proven itself in practical use. Our ‘ CoVRapid ... Study in Analytical and Bioanalytical Chemistry: Automated, flow-based chemiluminescence microarray immunoassay ...

Microarray Rapidly Identifies Antibodies Against SARS-CoV-2

"The system has already proven itself in practical use ... SARS-CoV-2 in human serum and plasma (CoVRapid CL-MIA).

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Analytical and Bioanalytical Chemistry,
13.05.2021 – DOI: 10.1007 ...

Versatile, fast and reliable: new automated microarray rapid test for detecting SARS-CoV-2 antibodies

So I started my A-Levels and chose subjects I ' d previously enjoyed and excelled in; Chemistry, Biology and History ... a more hands-on approach to learning and I ' m incredibly practical – I would often ...

Women in Engineering: An Interview with Engineering Apprentice, Morgan Pearce

“ The system helps train students skills of garbage classification and practical waste treatment courses at school ... processing time that previous solutions failed to do.

“ The manual version had been ...

‘ Green ’ school promotes waste value
Dr. Lowry has written 21 pool and spa water

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chemistry books, published more than 175 technical articles, written 29 technical white papers, penned a monthly column for Pool & Spa News, and crafted ...

Lab Manual

About the Book: The manual has been thoroughly revised, several new experiments and tests have been added while some redundant material has been deleted. Chapter 2 has been completely rewritten. An obvious change of this edition constitutes the splitting of Chapter 7 into two separate Chapters. Tables on derivatives of organic compounds have been expended. Also included are 20 estimations, 75 preparations and isolation experiments and

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approximately 135 in-text questions related to the experiments. The approximation of modern spectroscopic techniques to structure determination have been discussed in the last Chapter. This book is designed both for undergraduate and postgraduate level students with its enhanced and comprehensive presentation. This is an indispensable book for organic chemistry practicals. About the Author: Dr. Raj K. Bansal received his M.S. from the University of California, Davis, Calif, U.S.A., and Ph.D. from Calgary University, Calgary, Alberta, Canada. He was a postdoctoral fellow at the National Research Council (N.R.C.) of Canada in Halifax, N.S., Canada, followed by a Research Associateship at the Mellon Institute of Science, Carnegie-Mellon University, Pittsburgh Pa., U.S.A. Dr. Bansal has published a number of research papers in various foreign and Indian scientific

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journals. He is the author of six books on chemistry including this work-A Textbook of Organic Chemistry (5th ed., 2007), Organic Chemistry-Problems and Solutions (2nd edn., 2006), and Heterocyclic Chemistry (4th edn., 2005). One of his books, Synthetic Approaches in Organic Chemistry has been reprinted by Jones and Bartlett Publishers, Sudbury, Massachusetts, U.S.A. Dr. Bansal was a former Professor, Department of Chemistry, Indian Institute of Technology, Delhi, Hauz Khas, New Delhi.

Green chemistry involves designing novel ways to create and synthesize products and implement processes that will eliminate or greatly reduce negative environmental impacts. The Green Chemistry Laboratory Manual for General Chemistry provides educational laboratory materials that challenge students with the customary topics

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found in a general chemistry laboratory manual, while encouraging them to investigate the practice of green chemistry. Following a consistent format, each lab experiment begins with objectives and prelab questions highlighting important issues that must be understood prior to getting started. This is followed by detailed step-by-step procedures for performing the experiments. Students report specific results in sections designated for data, observations, and calculations. Once each experiment is completed, analysis questions test students' comprehension of the results. Additional questions encourage inquiry-based investigations and further research about how green chemistry principles compare with traditional, more hazardous experimental methods. By placing the learned concepts within the larger context of green chemistry principles, the lab manual enables students to see how these principles

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can be applied to real-world issues. Performing laboratory exercises through green experiments results in a safer learning environment, limits the quantity of hazardous waste generated, and reduces the cost for chemicals and waste disposal. Students using this manual will gain a greater appreciation for green chemistry principles and the possibilities for future use in their chosen careers.

The purpose of this "Manual of Practical Organic and Medicinal Chemistry" is intended for the Pharmacy students of D.Pharm, B.Pharm and Pharm.D students as per the new regulations of PCI-2016. This is specifically written to meet the present needs of revised curriculum.

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The edition of Comprehensive Practical Manual of Pharmaceutical Chemistry is authored in simple and comprehensive style according to PCI (Pharmacy Council of India) syllabus to meet the specific needs of the pharmacy students. It provides comprehensive yet concise chemistry for D.Pharmacy, B.Pharmacy, M.Pharmacy and Pharm D students. The main objective of this manual is to attract students to learn the basic theories of pharmaceutical chemistry thus the manual is aimed to enrich the inadequacy in teaching and learning of pharmaceutical chemistry by providing enormous information. The style of presentation of this manual is such that it not only gives deeper understanding of the subject but also will help the beginners to overcome the fright of the subject. The manual gives concise and pointwise information required during practicals in

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single book and eliminates the need of too many reference books during practicals. The manual authored in simple, lucid and easy language.

1. Basic Laboratory Techniques
1. To cut a glass tube or glass rod, 2. To bend the glass rod at an angle, 3. To draw a glass jet from a glass tube 4. To bore a cork and fit a glass tube into it Viva-Voce

2. Characterisation and Purification of Chemical Substances

1. To determine the melting point of the given unknown organic compound and its identification (simple laboratory technique)

Viva-Voce 2. To determine the boiling point of a given liquid when available in small quantity (simple laboratory method) Viva-

Voce 3. To prepare crystals of pure potash alum [$K_2SO_4 \cdot Al_2(SO_4)_3 \cdot 24H_2O$] from the given impure sample 4. To prepare the pure crystals of copper sulphate from the given

crude sample 5. To prepare pure crystals of

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benzoic acid from a given impure sample
Viva-Voce 3.Measurement of pH Values
1.To determine the pH value of vegetable juices, fruit juices, tap water and washing soda by using universal pH paper 2.To determine and compare the pH values of solutions of strong acid (HCl) and weak acid (CH₃COOH) of same concentration 3.To study the pH change in the titration of strong base Vs. strong acid by using universal indicator paper 4.To study the pH change by common ion (CH₃COO⁻ ion) in case of weak acid (CH₃COOH) 5.To determine the change in pH value of weak base (NH₄OH) in presence of a common ion (NH₄⁺) Viva-Voce 4.Chemical Equilibrium 1 To study the shift in equilibrium between ferric ions and thiocyanate ions by changing the concentrations of either of the ions 2.To study the shift in equilibrium between [Co(H₂O)₆]²⁺ and Cl⁻ ions by changing

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the concentrations of either of the ions Viva-Voce

5. Quantitative Analysis

1. To prepare M/10 oxalic acid solution by direct weighing method
2. To prepare M/10 solution of sodium carbonate by direct weighing method
3. To determine the strength of given solution of sodium hydroxide by titrating it against N/10 or M/20 solution of oxalic acid
4. To determine the strength of a given solution of hydrochloric acid by titrating it against a standard N/10 or M/20 sodium carbonate solution

Viva-Voce

6. Qualitative Analysis

Analysis of Anions

Analysis of Cations Viva-Voce

7. Detection of Elements in Organic Compounds

1. To detect the presence of nitrogen, sulphur and halogens in a given organic compound by Lassaigne ' s test
2. To detect the presence of nitrogen, sulphur and halogens in the given organic compound sample number by Lassaigne ' s test

Viva-Voce

INVESTIGATORY PROJECTS

1. Checking

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of Bacterial Contamination in Water 1. To check the bacterial contamination in drinking water by testing sulphide ions Viva-Voce 2. Methods of Water Purification 1. To purify water from suspended impurities by using sedimentation 2. To purify water by boiling 3. o purify water by distillation method 4. To purify water by reverse osmosis technique 5. To purify water by GAC method 6. To purify water by bleach treatment 7. To purify water by oxidising agent 8. To purify water by ozone treatment method Viva-Voce 3. Water Analysis 1. To test the hardness of different water samples Viva-Voce 4. Foaming Capacity of Various Soaps 1 .To compare the foaming capacity of different washing soaps 2. To study the effect of addition of sodium carbonate on foaming capacity of washing soap Viva-Voce 5. Tea Analysis 1. To study the acidity of different samples of tea leaves (tea) by using pH paper Viva-Voce 6. Analysis of

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Fruits and Vegetable Juices 1. To analysis the fruit and vegetable juices for the constituent present in them Viva-Voce 7. Rate of Evaporation 1. To study the rate of evaporation of different liquids IViva-Voce 8. Effect of Acids and Bases on Tensile Strength of Fibres 1. To compare the tensile strength of natural fibres and synthetic fibres 2. To study the effect of acids and bases on tensile strength of different fibres Viva-Voce

An Excellent Book in Accordance with the latest syllabus for Class-11 Prescribed by CBSE/NCERT and Adopted by Various State Education Boards. (A) Basic Laboratory Techniques – 1. To cut a glass tube or glass rod, 2. To bend the glass rod at an angle, 3. To draw a glass jet from a glass tube, 4. To bore a cork and fit a glass tube into it. (B) Characterisation and Purification of Chemical Substances- 1. To determine the melting point of the given unknown

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organic compound and its identification (simple laboratory technique), 2. To determine the boiling point of a given liquid when available in small quantity (simple laboratory method), 3. To prepare crystals of pure potash alum $[K_2SO_4 \cdot Al_2(SO_4)_3 \cdot 24H_2O]$ from the given impure sample, 4. To prepare the pure crystals of copper sulphate from the given crude sample, 5. To prepare pure crystals of benzoic acid from a given impure sample.

(C) Measurement of pH Values 1. To determine the pH value of vegetable juices, fruit juices, tap water and washing soda by using universal pH paper, 2. To determine and compare the pH values of solutions of strong acid (HCl) and weak acid (CH_3COOH) of same concentration, 3. To study the pH change in the titration of strong base Vs. strong acid by using universal indicator paper, 4. To study the pH change by common ion (CH_3COO^- –

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ion) in case of weak acid (CH_3COOH), 5. To determine the change in pH value of weak base (NH_4OH) in presence of a common ion (NH_4^+), (D) Chemical Equilibrium 1. To study the shift in equilibrium between ferric ions and thiocyanate ions by changing the concentrations of either of the ions, 2. To study the shift in equilibrium between $[\text{Co}(\text{H}_2\text{O})_6]^{2+}$ and Cl^- ions by changing the concentrations of either of the ions, (E) Quantitative Analysis 1. To prepare M/10 oxalic acid solution by direct weighing method, 2. To prepare M/10 solution of sodium carbonate by direct weighing method, 3. To determine the strength of given solution of sodium hydroxide by titrating it against N/10 or M/20 solution of oxalic acid, 4. To determine the strength of a given solution of hydrochloric acid by titrating it against a standard N/10 or M/20 sodium carbonate solution, (F) Qualitative

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Analysis 1. Analysis of Anions, 2. Analysis of Cations (G) Detection of Elements in Organic Compounds 1.To detect the presence of nitrogen, sulphur and halogens in a given organic compound by Lassaigne ' s test, 2. To detect the presence of nitrogen, sulphur and halogens in the given organic compound sample number by Lassaigne ' s test

INVESTIGATORY PROJECTS (A)

Checking of Bacterial Contamination in

Water 1.To check the bacterial

contamination in drinking water by testing

sulphide ions (B) Methods of Water

Purification 1.To purify water from

suspended impurities by using

sedimentation, 2. To purify water by boiling,

3.To purify water by distillation method,

4.To purify water by reverse osmosis

technique. 5.To purify water by GAC

method, 6.To purify water by bleach

treatment, 7.To purify water by oxidising

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agent, 8. To purify water by ozone treatment method. (C) Water Analysis 1. To test the hardness of different water samples. (D) Foaming Capacity of Various Soaps 1. To compare the foaming capacity of different washing soaps, 2. To study the effect of addition of sodium carbonate on foaming capacity of washing soap (E) Tea Analysis 1. To study the acidity of different samples of tea leaves (tea) by using pH paper (F) Analysis of Fruits and Vegetable Juices 1. To analyse the fruit and vegetable juices for the constituent present in them (G) Rate of Evaporation 1. To study the rate of evaporation of different liquids (H) Effect of Acids and Bases on Tensile Strength of Fibres 1. To compare the tensile strength of natural fibres and synthetic fibres, 2. To study the effect of acids and bases on tensile strength of different fibres. Log & Antilog Table

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