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VS Electron

Geometry - The Effect
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Chapter 9 Molecular
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12. The Shapes of
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For this activity, you will need to refer to your Lewis Dot Worksheet # 1.

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9. Explain the

difference between a

bonding electron

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domain and a nonbonding electron domain using the examples in Model 1.

10. Circle the correct word or phrase to complete the sentences: Pairs of electrons will (attract/repel ...

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polyatomic ions are not all flat structures. Each has a three-dimensional shape that helps account for its various chemical and physical properties. [9-12 Content Standard B - Structure and properties of matter] Students often find it difficult to

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The molecular geometry main shapes are tetrahedral, trigonal planar, trigonal pyramidal, bent, and linear and are named by measuring the bond angles between the central atom and another atom bonded

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to it. Molecular

Geometry

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~~Worksheet & Lab~~

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In this activity you will learn how to predict molecular shapes.

Model 1 □ Lewis

Structures Lone pair =

□□ H 2 CO 3 electron

domains (3 bonding, 0

nonbonding) 3-D

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Molecular Shape BeF

2 2 electron domains
(2 bonding, 0

nonbonding) CH 4 4
electron domains (4
bonding, 0

nonbonding) NH 3 4
electron domains (3
bonding, 1

nonbonding) H 2 O 4
electron domains

~~20 Molecular~~

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Schow's Chemistry Classes

electron pair

geometry: molecular

geometry: shape:

polarity: BeH 2: 2: 0:

linear: MX 2: linear:

nonpolar: BH 3: 3: 0:

trigonal planar: MX 3:

trigonal planar:

nonpolar: SO 2: 2: 1:

trigonal planar: MX 2

E: bent: polar : CH 4:

4: 0: tetrahedral: MX

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4: tetrahedral:

nonpolar: NH₃: 3: 3: 1:

tetrahedral : MX₃ E:

trigonal . pyramidal.

polar: H₂O: 2: 2:

tetrahedral: MX₂ E₂:

bent or . V-shaped.

polar: PCl₅: 5: 5: 0:

trigonal . bipyramidal.

MX₅: trigonal .

bipyramidal. nonpolar:

SF₆: 6: 6: 0

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How to Determine
Molecular Geometry □

YouTube: This video describes one method for quickly finding the major geometrical shapes for simple molecules. Molecular Geometries The VSEPR theory describes five main shapes of simple molecules: linear,

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trigonal planar,
tetrahedral, trigonal
bipyramidal, and
octahedral.

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Building Molecules □

Molecular Geometry

ACTIVITY: VERSION

A (Advanced

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will be able to□. · Draw

the Lewis Dot

Structure for

molecules. · Identify

the number of Bonded

Atoms and Lone Pairs

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on a central atom. •

Identify the bond
angle in molecules.

Activity Answers

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