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Section 9.4 □ Naming and Writing Formulas for Acids and Bases. An acid is a compound that produces  $H^+$  ions when it dissolves in water. The formula for an acid normally starts with an H. When naming acids, you should first determine whether or not the acid contains oxygen.

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Chapter 9 Chemistry Chemical Names and Formulas. Acids. base. Cation. Ionic compounds. Compounds that contain one or more hydrogen atoms and produce $\square$ . An ionic compound that produces hydroxide ions when dissolved $\square$ . Any atom or group of atoms that has a positive charge. Compounds composed of cations and anions.

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Chapter 9 Chemical Names and Formulas83 SECTION 9.3 NAMING AND WRITING FORMULAS FOR MOLECULAR COMPOUNDS (pages 268 $\square$ 270) This section explains the rules for naming and writing formulas for binary molecular compounds. Naming Binary Molecular Compounds (pages 268 $\square$ 269) 1. Circle the letter of the type(s) of elements that form binary molecular compounds.

## ~~Name Date Class CHEMICAL NAMES AND FORMULAS 9~~

SECTION 9.2 NAMING AND WRITING FORMULAS FOR IONIC COMPOUNDS I. Write the

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formulas for these binary ionic compounds. c. potassium iodide 2. write the formulas for the compounds formed c. b. f. 3. Name the following binary ionic compounds. sodium sulfide g. CuCl: h.  $\text{SnCl}_2$ . a.  $\text{MnO}_2$  c. d.  $\text{SrBr}_2$  221 Chapter 9 Chemical Names and Formulas

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Chapter 9 Chemical Names and Formulas 9.1 Naming Ions 9.2 Naming and Writing Formulas for Ionic Compounds 9.3 Naming and Writing Formulas for Molecular Compounds 9.4 Naming and Writing Formulas for Acids and Bases 9.5 The Laws Governing How Compounds Form

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Chapter 9 - Chemical Names and Formulas ... in parenthesis after the name of the metal (Table 9.2, p.255) Predicting Ionic Charges Some of the post-transition elements also have more than one possible oxidation state. Tin (II) =  $\text{Sn}^{2+}$  Lead (II) =  $\text{Pb}^{2+}$  Tin (IV) =  $\text{Sn}^{4+}$  Lead (IV) =  $\text{Pb}^{4+}$

~~- Chemical Names and Formulas -~~

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Chemical Names and Formulas 281 CHAPTER 9 Assessment 42. a. 2<sup>-</sup> b. 1<sup>+</sup> c. 1<sup>-</sup> d. 3<sup>+</sup> 43. a. 2<sup>+</sup> b. 2<sup>+</sup> c. 3<sup>+</sup> d. 1<sup>+</sup> 44. a. barium ion b. iodide ion c. silver ion d. mercury(II) ion 45. cyanide, CN<sup>-</sup> and hydroxide, OH<sup>-</sup> 46. a. hydroxide ion b. lead(IV) ion c. sulfate ion d. oxide ion 47. zero; A compound is electrically neutral. 48. The symbols for the cation and

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Chapter 9 Practice Problems: Chemical Names and Formulas. Section 9.1: Naming Ions. 1. What is the charge on the ion typically formed by each element? a. oxygen c. sodium e. nickel, two electrons lost . b. iodine d. aluminum f. magnesium. 2. How many electrons does the neutral atom gain or lose when each ion forms? a. Cr<sup>3+</sup> c. Li<sup>+</sup> e. Cl<sup>-</sup> b. P<sup>3-</sup> ...

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