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||PART 1 BY SANJIV SIR ~~Chapter 4 Atomic Structure~~
4.2 The Structure of an Atom We have learnt Dalton ' s
atomic theory in Chapter 3, which suggested that the atom
was indivisible and indestructible. But the discovery of two
fundamental particles (electrons and protons) inside the
atom, led to the failure of this aspect of Dalton ' s atomic

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theory. It was then considered necessary to

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Chapter 4: The Structure of the Atom. 100. The Structure of the Atom. Graphite surface. BIGIdeaAtoms are the

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fundamental building blocks of matter. 4.1 Early Ideas About Matter. MAIN Idea The ancient Greeks tried to explain matter, but the scientific study of the atom began with John Dalton in the early 1800s.

~~Chapter 4: The Structure of the Atom – Mrs. Taylor's Classes~~

1. All elements are composed of tiny invisible particles (atoms) 2. All atoms of the same element are identical. 3. Atoms of different elements can physically combine together or chemically combine in simple whole number ratios to form compounds. 4. Chemical reactions occur when atoms are separated, joined, or rearranged.

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Multiple Choice, Fill in the Blank, Short Answer Directions:

Read all questions carefully and answer to the best of your ability. Do not leave any questions blank! Show your work where needed

1. How long ago did Democritus propose the idea of atoms?

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1. CHAPTER - 4 STRUCTURE OF THE ATOM CLASS :- IX

MADE BY :- SHIVA PRASAD SCHOOL :- K.V. 2, SADRAS ,

KALPAKKAM

2. 1a) Charged particles in matter :- Atoms

have three types of sub atomic particles. They are electrons, protons and neutrons. Electrons are negatively charged (e^-),

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protons are positively charged (p^+) and neutrons have no charge (n).

~~structure of the atom - SlideShare~~

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Atomic structure refers to the structure of atom comprising a

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nucleus (center) in which the protons (positively charged) and neutrons (neutral) are present. The negatively charged particles called electrons revolve around the center of the nucleus.. The history of atomic structure and quantum mechanics dates back to the times of Democritus, the man who first proposed that matter is composed of ...

~~Atomic Structure – Electrons, Protons, Neutrons and Atomic~~

...

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Example Beryllium : atomic number 4, mass number 9. It has 4 protons, 4 electrons and $9-4 = 5$ neutrons Isotopes Atoms of the same element can have different numbers of neutrons; these atoms are called isotopes of that element. Isotopes of an element have the same chemical properties because they have the same electronic structure

~~Revision Guide: 4.1 Atomic Structure and the Periodic Table~~

4.05 Atomic Number, Mass Number, Isotopes and Isobars

4.05 Atomic Number, Mass number, Isotopes and Isobars.

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(4.1) Dalton's atomic theory 1. all elements are made of tiny indivisible particles called atoms 2. atoms of the same element are identical and those of different elements are different 3. atoms of different elements can physically mix or chemically combine together in whole number ratios to form compounds

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