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Answers Chapter 19 - Acids, Bases, and Salts - 19 Assessment - Page 684: 52 Answer The Arrhenius definition states that acids are compounds that contain hydrogen, and when placed in aqueous solution, ionize to form hydrogen (H^+)

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An acid is a substance that contains hydrogen and ionizes to produce hydrogen ions in aqueous solution; A base is a substance that contains a hydroxide group and dissociates to produce a hydroxide ion in aqueous solution. Bronsted-Lowry model. A model of acids and bases in which an acid is a hydrogen-ion donor and a base is a hydrogen-ion acceptor.

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states that acids are compounds that contain hydrogen, and when placed in aqueous solution, ionize to form hydrogen (H^+ ions).

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two substances that are related by the loss or gain of a single hydrogen ion; ammonia (NH_3) and the ammonium ion (NH_4^+) are a conjugate acid-base pair (19.1) monoprotic acid. any acid that contains one ionizable proton (hydrogen ion); nitric acid (HNO_3) is a monoprotic acid (19.1) diprotic acid.

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596 Chapter 19 Acids and Bases PRACTICE PROBLEMS 1. Write a balanced formula equation for the reaction that occurs between each of the following pairs of reactants. a. magnesium and nitric acid b.

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aluminum and sulfuric acid c. calcium carbonate and hydrobromic acid d. potassium hydrogen carbonate and hydrochloric acid

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