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equivalence point must be reached., 2.

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acids or bases act as acids or bases, they
form their conjugate and the strength is the
inverse of their parent. and there are
certain strong acids and strong bases. LiOH ,
 NaOH , KOH - when act as bases (get rid of H^+
ion) then left with Li Page 13/24

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ACIDS AND BASES 453 SECTION 15-1 0 OBJECTIVES

List five general properties of aqueous acids and bases. Name common binary acids and oxyacids, given their chemical formulas. List five acids commonly used in industry and the laboratory, and give two properties of each. Define acid and base according to Arrhenius's theory of ionization. Explain the differences

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Questions 1) _____ is found in carbonated
beverages due to the reaction of carbon
dioxide with water. Chapter 15 Acids and
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Chapter 15 Acids and Bases. strong acid.
strong base. weak acid. weak base. an acid
that ionizes

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Chapter 15: Acids and Bases Acids and Bases
Arrhenius Definitions: acids - compounds that produce an increase in $[H^+]$ when dissolved in water bases - compounds that produce an

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increase in $[OH^-]$ when dissolved in water

Lewis Definitions: acids - electron pair

acceptors bases - electron pair donors

Brønsted-Lowry Definitions:

Chapter 15: Acids and Bases Acids and Bases
Chapter 15 Acids Bases Review bases C. They
can act as acids or bases when reacting with
each other D. They can act as acids when
reacting with HCl molecules 14) 25.0 cm³ of
sulfuric acid solution reacts with 36.2 cm³
of 0.225 mol dm⁻³ sodium hydroxide solution.
The concentration of the Chapter 15 Acids

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Eventually, you will agreed discover a extra
experience and talent by spending more cash.
still when? complete you consent that you
require to acquire those every needs past
having significantly cash?

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bitter L2. pH B. ionize only slightly in solution £3. titration C. salt + H₂ (g) F 4. neutralization D. salt +...

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strong acid. weak acid. strong base. weak
base. an acid that ionizes completely in a
solvent. an acid that releases few hydrogen
ions in aqueous solution. a base that ionizes
completely in a solvent. a base that releases
few hydroxide ions in aqueous solution.

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