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questionbase.50megs.com AS-Level Revision Notes AS Chemistry – Revision Notes Unit 1 – Atomic Structure, Bonding And Periodicity Atomic Structure 1. All atoms have a mass number, A (the number of nucleons), and a proton number, Z (the number of protons). 2.

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Revision for AQA Chemistry AS and A-Level from 2015, including summary notes, worksheets and past exam questions for each topic and paper.

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These revision notes include all of the content for Chemistry paper 1 from the new AQA specification. They include some examples for the more mathematical topics as well as descriptions for all the Required Practicals included on Paper 1. The notes are 53 pages long, but could be printed 2 or 4 to a page to make them into a shorter booklet for pupils to revise from.

AQA GCSE Chemistry Paper 1 Revision notes | Teaching Resources

Revision for AQA Chemistry GCSE, including summary notes, exam questions by topic and videos for each module

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Chemistry is the study of the composition, behaviour and properties of matter, and of the elements of the Earth and its atmosphere.

GCSE Chemistry (Single Science) - BBC Bitesize

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To aid with revision we suggest that pupils use the topic summary notes to help them attempt the problem sheets for every topic. The problem sheets contain a mixture of multiple choice and short answer questions, many of which have been taken from past papers. Unit 1 Summary Notes Unit 1 Summary Notes

Higher Chemistry - PRESTWICK ACADEMY

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National 4/5 Chemistry Revision Notes . Unit 1: Chemical Changes and Structure . Subsection 1: SUBSTANCES . 1. Element: is a substance that cannot be broken down into anything simpler. An element is made of the same type of atom. 2. Compound: is formed when 2 or more elements chemically join. e.g. sodium chloride . 3. Naming compounds

National 4/5 Chemistry Revision Notes

Chemistry GCE A Level Revision Guide (2015 specification) – Units 1 and 2. Chemistry . KS5 > Revision booklets prepared by Bangor University for the 2015 specification. Science. Revision. A level. Unit 1 and 2. Files. Open resource. Feedback. Not seeing what you want? Is there are problem with the files? Do you have a suggestion? Please give ...

Resource WJEC Educational Resources Website

IGCSE Chemistry revision notes made for the CIE exam boards. This covers all the topics and modules for all specifications including 0620. Topics that

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are covered include 1. The particulate nature of matter; 2.

CIE IGCSE Notes | A* Chemistry

Na = 23, O = 16, H = 1. (You will be given these numbers in the exam.) Next, divide the atomic number of oxygen by the total atomic number of NaOH, and multiply by 100 to get a percentage. NaOH is $23 + 16 + 1 = 40$. $(16 \div 40) \times 100 = 0.4 \times 100 = 40\%$. So the percentage by mass of oxygen in sodium hydroxide is 40%.

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