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Question 11. Enthalpy of combustion of carbon to carbon dioxide is $-393.5 \text{ J mol}^{-1}$. Calculate the heat released upon formation of 35.2 g of CO_2 from carbon and oxygen gas. Answer: The combustion equation is: $\text{C (s)} + 0.2 \text{ (g)} \rightarrow \text{CO}_2 \text{ (g)}$; $\Delta H = -393.5 \text{ KJ mol}^{-1}$. Heat released in the formation of 44g of $\text{CO}_2 = 393.5 \text{ kJ}$.

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